

Read Text for Comprehension

Read this article for deeper understanding. No summary is required, although you may want to circle, underline, or mark key ideas and words.

Some of the examples of ionic compounds are listed below:

1. Magnesium chloride $\rightarrow MgCl_2$
2. Barium fluoride $\rightarrow BaF_2$
3. Calcium phosphate $\rightarrow Ca_3(PO_4)_2$
4. Lithium phosphate $\rightarrow Li_3PO_4$
5. Sodium nitride $\rightarrow Na_3N$
6. Aluminum phosphide $\rightarrow AlP$
7. Aluminum nitride $\rightarrow AlN$
8. Barium oxide $\rightarrow BaO$
9. Calcium bromide $\rightarrow CaBr_2$
10. Sodium fluoride $\rightarrow NaF$
11. Sodium hydroxide $\rightarrow NaOH$
12. Calcium chloride $\rightarrow CaCl_2$
13. Calcium hydroxide $\rightarrow Ca(OH)_2$
14. Potassium iodide $\rightarrow KI$
15. Sodium sulfide $\rightarrow Na_2S$
16. Aluminum sulfide $\rightarrow Al_2S_3$
17. Potassium sulfide $\rightarrow K_2S$
18. Lithium nitrate $\rightarrow LiNO_3$
19. Sodium carbonate $\rightarrow Na_2CO_3$
20. Strontium fluoride $\rightarrow SrF_2$
21. Sodium iodide $\rightarrow NaI$
22. Aluminum bromide $\rightarrow AlBr_3$
23. Calcium oxide $\rightarrow CaO$
24. Magnesium sulfide $\rightarrow MgS$
25. Calcium nitrate $\rightarrow Ca(NO_3)_2$
26. Sodium hydroxide $\rightarrow NaOH$
27. Magnesium sulfate $\rightarrow MgSO_4$

So we can say that ionic compounds are formed by combination of cation and anions which are bonded through electrostatic force of attraction. The charges of ions are always balance with opposite charges and form a neutral ionic compound.