

Topic Introduction



Summarize your understanding of each paragraph.

The limiting reagent (or limiting reactant, LR) in a chemical reaction is the substance that is totally consumed when the chemical reaction is complete. The amount of product formed is limited by this reagent, since the reaction cannot continue without it.

Percent yield = actual yield / theoretical yield. The actual yield is a product that is obtained by experimentation. The theoretical yield is obtained through stoichiometric calculation. If the two yields are equal, you have 100 % yield. Usually you obtain less than 100 %.

Percentage yield is important because: chemical reactions very often form by-products as well as the intended product. In most reactions, not all of the reactants actually react.

Remember, the calculated or expected amount of product is called the theoretical yield. The amount of product actually produced is called the actual yield. When you divide actual yield by theoretical yield you get a decimal percentage known as the percent yield.

Read/Summarize Text



1. Read the passage.
2. Underline key expressions in each sentence.
3. Re-write each word (or expression) you underlined.
4. Summarize the passage.

Title of Passage.

1

Calculating Theoretical Yield

1. Identify your desired product.
2. Write down the number of moles of your limiting reactant.
3. Find the ratio of molecules in your product and reactant.
4. Multiply the ratio by the reactant's quantity in moles.
5. Convert the result to grams.

2

<https://www.wikihow.com/Calculate-Percent-Yield-in-Chemistry>

Re-write words you underlined

3

Using a complete sentence, summarize or rephrase the passage

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