

Topic Introduction



Summarize your understanding of each paragraph.

Energy levels inside an atom are the specific energies that electrons have when occupying specific orbitals. Electrons can be excited to higher energy levels by absorbing energy from the surroundings.

Light is emitted when an electron relaxes from a high energy state to a lower one.

The energy of an electron in an atom is not continuous, but quantized. The energies corresponding to each of the allowed orbitals are called energy levels.

An electron can't have an energy value half way between two energy levels in an atom.

Read/Summarize Text



1. Read the passage.
2. Underline key expressions in each sentence.
3. Re-write each word (or expression) you underlined.
4. Summarize the passage.

Atoms.

Atoms are submicroscopic particles that make up the matter we see around us. Some 92 kinds of atoms exist in nature, such as hydrogen, oxygen, iron, gold, etc. That is, gold is made of gold atoms, and iron of iron atoms. Other substances are made of fixed combinations of atoms, called molecules. For example, water is made of two hydrogen atoms linked to an oxygen atom, hence its chemical formula, H₂O. Atoms are themselves made up of even smaller particles. Each atom consists of a core, or more technically a nucleus around which yet smaller particles called electrons orbit. The nucleus is itself composed of smaller particles called protons and neutrons. Protons, neutrons, and electrons differ in several ways.

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Re-write words you underlined

3

Using a complete sentence, summarize or rephrase the passage

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